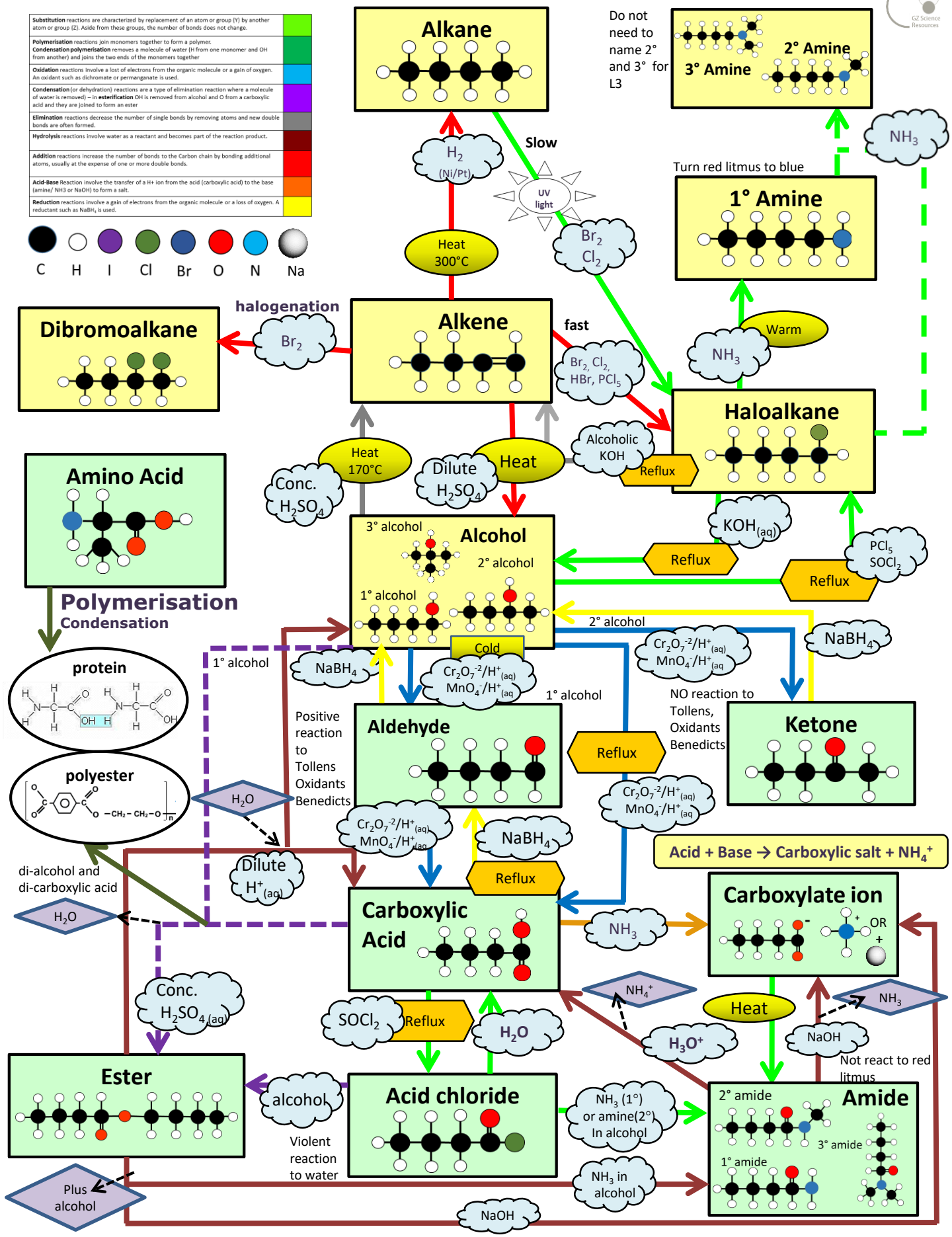
NCEA L3 Organic Chemistry Organic Scheme

Substitution reactions are characterized by replacement of an atom or group (Y) by another atom or group (Z). Aside from these groups, the number of bonds does not change.	
Polymerisation reactions join monomers together to form a polymer. **Condensation polymerisation** removes a molecule of water (H from one monomer and OH from another) and joins the two ends of the monomers together	
Oxidation reactions involve a lost of electrons from the organic molecule or a gain of oxygen. An oxidant such as dichromate or permanganate is used.	
Condensation (or dehydration) reactions are a type of elimination reaction where a molecule of water is removed) – in **esterification** OH is removed from alcohol and O from a carboxylic acid and they are joined to form an ester	
Elimination reactions decrease the number of single bonds by removing atoms and new double bonds are often formed.	
Hydrolysis reactions involve water as a reactant and becomes part of the reaction product.	
Addition reactions increase the number of bonds to the Carbon chain by bonding additional atoms, usually at the expense of one or more double bonds.	
Acid-Base Reaction involve the transfer of a H+ ion from the acid (carboxylic acid) to the base (amine/ NH3 or NaOH) to form a salt.	
Reduction reactions involve a gain of electrons from the organic molecule or a loss of oxygen. A reductant such as $NaBH_4$ is used.	

C H I Cl Br O N Na