**Chemistry 2.6 AS 91166** Demonstrate understanding of chemical reactivity



Writing Excellence answers to **Equilibrium – Pressure** questions

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| **Equilibrium – Pressure QUESTION** |
| **Question:**  The two reactions shown in the following table are both at equilibrium.Compare and contrast the effect of increasing the pressure on both reactions, with reference to the equilibrium positions. |
| **ANSWER** |
| 1. State the **equilibrium principle** |  |
| 2. Describe the **factor** in your question AND Link increasing the principle to how the **system responds**[some questions will be decreasing] |  |
| 3. **Generally,** explain which side of the equation is favoured (relate to moles) AND the general observations – at visible and particle level. |  |
| 4. **Specifically,** in reaction one describe number of moles in both sides of the equation AND link to which direction of reaction would be favoured (and observation) |  |
| 5. **Specifically**, in reaction two link number of moles in both sides of the equation to observation AND link to which direction of reaction would be favoured  |  |
| 6. Describe how the **system shift** in reaction two would effect at particle level AND final observation.  |  |

NOTE: The white column is how your answer would appear on your test paper so make sure you **write out complete sentences**. The grey area is just to help you structure your answer and would not appear in the question.