**Chemistry 3.6 AS 91392** Demonstrate understanding of equilibrium principles in aqueous systems



Writing Excellence answers to **Titration Curve – After the Start pH** questions

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| **Titration Curve – after the Start pH QUESTION**  |
| **Question:**   **:** 20.0 mL of 0.0896 mol L–1 ethanoic acid is titrated with 0.100 mol L–1 sodium hydroxide. p*K*a (CH3COOH) = 4.76Calculate the pH of the titration mixture after 5.00 mL of NaOH has been added. Kw = 1 x 10-14  |
| **ANSWER** |
| **1.** determine if starting solution is acid or base (will it accept or donate H+) – strong or weakAnd write down all available information |  |
| **2.** Write down neutralisation equation |  |
| **3.** calculate  *n*(CH3COOH at start) n = cv*3sgf and units*  |  |
| **4.** calculate *n*(NaOH) and therefore n(CH3COO-)n = cv*assume n*(NaOH) =  *n*(CH3COO-)*3sgf and units* |  |
| **5.** calculate  *n*(CH3COOH) After 5 mL NaOH added: (total 25mL)=(n(CH3COOH – n(CH3COO-) after 5mL)*3sgf and units* |  |
| **6**. calculate c(CH3COOH)c = n/v*3sgf and units* |  |
| **7**. calculate [H3O+] Ka = 10-pKa[H3O+] = √ Ka x c(HA)*3sgf and units* |  |
| **8**. Calculate pHpH = -log [H3O+] *3sgf* *Check pH against estimate on curve* |  |

NOTE: The white column is how your answer would appear on your test paper so make sure you **write out complete sentences**. The grey area is just to help you structure your answer and would not appear in the question.